Name_____

Hour_____

Section 2-1/2-2 Review

A. Use the vocabulary terms in the following list, fill in the blanks in the statements below.

atom	**atomic number**	**compound**
electron	**element**	**mass number**
neutron	**nucleus**	**proton**

1. A substance that cannot be broken down into other substances by ordinary chemical means is a(n)_____.
2. A substance formed by the chemical combination of two or more elements is a(n)_____.
3. The basic unit of structure of all elements is the_____.
4. Atoms are made up of three types of particles:_____, _____, and _____.
5. The dense central portion of the atom is the_____.
6. The number of protons in the nucleus of its atom is the_____ of an element.
7. The number of protons plus the number of neutrons in the nucleus of an atom is its_____.

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colloid	**covalent bonding**	**ionic bonding**
isotope	**mixture**	**pH**
solution	**suspension**	

1. Different varieties of the same element having different numbers of neutrons in the nuclei are called_____.

9. Chemical bonding in which there is a transfer of electrons from one atom to another is_____.

10. Chemical bonding in which there is a sharing of electrons between atoms is_____.

11. Measurement of the hydrogen ion concentration of a solution may be given in terms of _____.

12. A purely physical association of substances is a_____.

13. Three types of mixtures are_____, _____, and_____.

B. If the statement is true, write **TRUE** in the space at the left. If the statement is false, write **FALSE** in the space at the left.

_____1. Compounds are the pure substances that are the building blocks of matter.

_____2. Only about 24 elements are commonly found in living organisms.

_____3. The mass of an atom is mainly in the neutron.

_____4. An electron has a positive charge.

_____5. The second energy level can hold up to two electrons.

_____6. A substance in which two or more different elements are chemically combined is called a compound.

_____7. Chemical bonding in which electrons are shared is called ionic.

_____8. Atoms that carry positive or negative charges are known as solids.

_____9. Particles are likely to be farthest apart in a liquid.

_____10. The freezing of water is a physical change.

_____11. In a mixture, there is no bonding of atoms.

_____12. A solute dissolves in a solvent.

_____13. Adding starch to water and stirring will form a solution.

_____14. Jello is an example of a colloidal suspension.

_____15. A solution with fewer OH^- ions than H^+ ions is a base.

C. Fill in the blanks with the correct vocabulary word.

1. The nucleus, the center of the atom, is made up of _____and _____.
2. The negatively charged particles in atoms are called_____.
3. Different isotopes of the same element have different numbers of _____.
4. In a(an)_____bond, electrons are transferred from one atom to another.

D. Match the following term with the best definition.

a. polarity b. acidic c. basic

_____1. unequal sharing of electrons

_____2. lemon juice, pH 1.5

_____3. lower concentrations of H+ ions than pure water

_____4. ammonia, pH 11.5

_____5. a slight negative charge at one end of a molecule, a slight positive charge at the other end

_____6. pH values that are below 7

_____7. alkaline solutions